

First Year Pharm. D : SUMMER - 2019
SUBJECT: PHARMACEUTICAL INORGANIC CHEMISTRY

Day: Wednesday
Date: 17/04/2019

Time: 10.00 A.M. TO 01.00 P.M.
Max. Marks: 70

S-2019-4502

N.B:

- 1) **Q.No.1 and Q.No.5 are COMPULSORY.** Out of remaining solve any **TWO** Questions from Section-I and Section-II
- 2) Answer to both the section should be written in **SEPARATE** answer book.
- 3) Figures to the right indicate **FULL** marks.

SECTION-I

- Q.1** A) Attempt any **FOUR** of the following. (08)
- i) Give the list of reagents required for limit test for Sulphate with use of each.
 - ii) List out the conditions that are responsible for Hyperkalemia.
 - iii) Give the classification of Volumetric analysis.
 - iv) Give the principle and reaction involved in the assay of Ammonium chloride.
 - v) Write biochemical role of Iron as essential and trace element.
- B) Give the properties of Radioactive substances. (03)
- Q.2** What do you mean by Redox reaction? Explain in detail Ion electron balance method for calculating equivalent weight. Give the procedure, principle and reaction involved in the Assay of Ferrous sulphate. (12)
- Q.3** a) Discuss in detail Neutralization curves for Acid-base titrations. (07)
b) Give the limit test for Chloride in detail. (05)
- Q.4** Write short note on any **THREE** of the following. (12)
- a) Salt hydrolysis
 - b) Oxidation number method
 - c) Limit test for Iron
 - d) Theories of Acid-base indicator

SECTION-II

- Q.5** A) Attempt any **FOUR** of the following. (08)
- i) Why Complexometric titrations are carried out in buffer solutions.
 - ii) Give the properties, uses and storage conditions of CO₂ gas.
 - iii) Give the mechanism of action of Saline Cathartics.
 - iv) What are the side effects associated with Antacid therapy?
 - v) Give the properties, uses and storage conditions of O₂ gas.
- B) Give the preparation and standardization of 0.1 M Perchloric acid. (03)
- Q.6** Discuss in detail the end point determination in Complexometric titrations. Explain types of EDTA titration with example. (12)
- Q.7** a) What do you mean by Non-aqueous titrations? Give the various types of Non-aqueous solvents used in Non-aqueous titrations. (07)
b) Explain filtration medias used in Gravimetric analysis in detail (05)
- Q.8** Write short note on any **THREE** of the following. (12)
- a) Dentifrices
 - b) Metalochrome indicators in Complexometric titrations.
 - c) Co-precipitation and Post-precipitation
 - d) Gastrointestinal Protectives and Adsorbents

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